

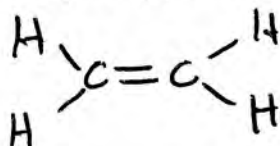
Suggested
NAME

Answers
SIGNATURE

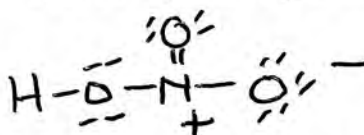
Drawing Lewis Structures

1) Write Lewis Structures for each of the following molecules. Include charges if needed (8 pts).

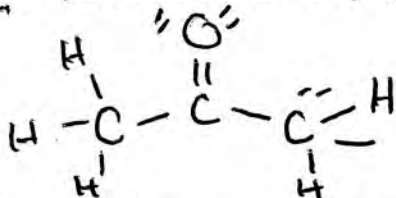
a) $\text{CH}_2=\text{CH}_2$ (ethene)



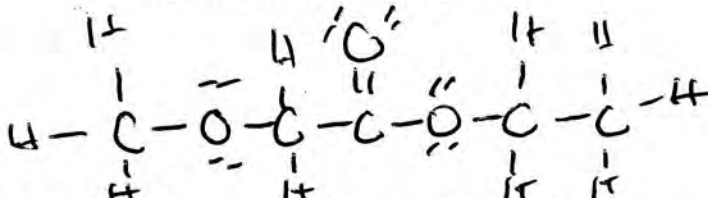
b) HONO_2 (HNO_3 , nitric acid)



c) acetone anion: $\text{CH}_3\text{C}(\text{O})\text{CH}_2^-$

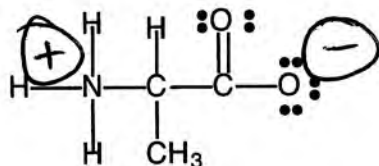


d) $\text{CH}_3\text{OCH}_2\text{CO}_2\text{CH}_2\text{CH}_3$

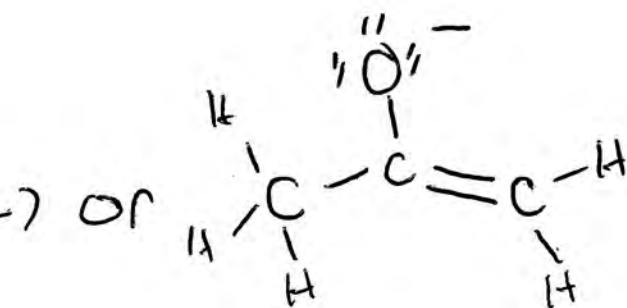


2) Add a formal charge add to each atom with non-zero charge to complete the following structures (2 pts).

a) alanine



b) fulminic acid



two resonance structures possible